

Chemical Equations Converted

 **SaitechAI**

Here is the combined table from your "Redox Reactions Worksheet" with all the reactions and their respective details filled in:

S. No.	Chemical Reaction	Substance Oxidized	Substance Reduced	Oxidizing Agent	Reducing Agent
1	$\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$	H_2	CuO	CuO	H_2
2	$\text{H}_2\text{S} + \text{Cl}_2 \rightarrow \text{S} + 2\text{HCl}$	H_2S	Cl_2	Cl_2	H_2S
3	$\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$	C	ZnO	ZnO	C
4	$\text{MnO}_2 + 4\text{HCl} \rightarrow \text{MnCl}_2 + 2\text{H}_2\text{O} + \text{Cl}_2$	HCl	MnO_2	MnO_2	HCl
5	$\text{SO}_2 + 2\text{HS} \rightarrow 2\text{H}_2\text{O} + 3\text{S}$	HS	SO_2	SO_2	HS
6	$\text{H}_2\text{O} + \text{F}_2 \rightarrow \text{HF} + \text{HOF}$	H_2O	F_2	F_2	H_2O
7	$4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$	NH_3	O_2	O_2	NH_3
8	$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$	CO	Fe_2O_3	Fe_2O_3	CO
9	$3\text{Fe} + 4\text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + 4\text{H}_2$	Fe	H_2O	H_2O	Fe
10	$\text{Pb}_3\text{O}_4 + 8\text{HCl} \rightarrow 3\text{PbCl}_2 + \text{Cl}_2 + 4\text{H}_2\text{O}$	HCl	Pb_3O_4	Pb_3O_4	

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Here's the table with the specified columns ("Substance Oxidized", "Substance Reduced", "Oxidizing Agent", and "Reducing Agent") left blank as per your request:

S. No.	Chemical Reaction	Substance Oxidized	Substance Reduced	Oxidizing Agent	Reducing Agent
1	$\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$				
2	$\text{H}_2\text{S} + \text{Cl}_2 \rightarrow \text{S} + 2\text{HCl}$				

S. No.	Chemical Reaction	Substance Oxidized	Substance Reduced	Oxidizing Agent	Reducing Agent
3	$\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$				
4	$\text{MnO}_2 + 4\text{HCl} \rightarrow \text{MnCl}_2 + 2\text{H}_2\text{O} + \text{Cl}_2$				
5	$\text{SO}_2 + 2\text{HS} \rightarrow 2\text{H}_2\text{O} + 3\text{S}$				
6	$\text{H}_2\text{O} + \text{F}_2 \rightarrow \text{HF} + \text{HOF}$				
7	$4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$				
8	$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$				
9	$3\text{Fe} + 4\text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + 4\text{H}_2$				
10	$\text{Pb}_3\text{O}_4 + 8\text{HCl} \rightarrow 3\text{PbCl}_2 + \text{Cl}_2 + 4\text{H}_2\text{O}$				