$n_1 = $ no. of moles of solvent

$n_2 = $ no. of moles of solute

	e
H_2SO_4	2
HCl	1
V_2O_5	
$KMnO_4$	1
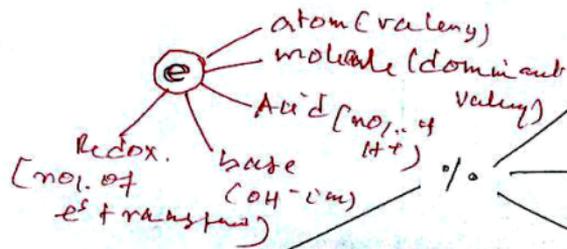
atom (valency)

molecule (dominant valency)

Acid [no. of H^+]

Redox [no. of e^- transfer]

base $(OH^- \text{ ion})$

$\dfrac{W_2}{W} \times 10^6 - $ p|pm parts per million

$w/w \% = \dfrac{w_2}{w} \times 100 \quad -g$

$w/v \% = \dfrac{w_2}{v} \times 100 \quad -g -ml$

$v/v \% = \dfrac{V_2 - ml}{V} \times 100 \quad -ml$

$\%$

C

$M = \dfrac{n_2}{v} = \dfrac{w_2}{M_2} \times \dfrac{1}{v} = \dfrac{w_2 \times 1000}{M_2 \quad v - L}$
Molarity $-gmol^{-1}$

mole fraction

X_2

$\dfrac{n_2}{n_1 + n_2} = \dfrac{n_2}{n}$

$N = \dfrac{eq_2}{v}$
Normality $= \dfrac{w_2}{GEW_2} \times \dfrac{1}{v} \quad -L$

m molality

$= \dfrac{n_2}{w} \quad -kg$

$= \dfrac{w_2}{M_2} \times \dfrac{1}{w}$

$= \dfrac{w_2}{M_2} \times \dfrac{1000}{w} \quad -g$ (dil. soln.)

$= \dfrac{w_2}{M_2/e} \quad v - L$

$= \dfrac{w_2}{M_2} \times e = Me$

valency factor

$w = $ weight of solution

$w_2 = $ weight of solute

$w_1 = $ weight of solvent

$= \dfrac{n_2}{n_1}$ (dil. soln.)

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