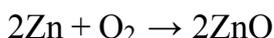


# Redox Reactions of d-block Elements



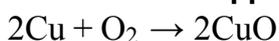
Here are ten chemical reactions involving d-block elements that showcase redox behavior:

**1. Oxidation of Zinc:**



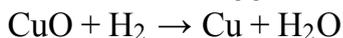
Zinc is oxidized to zinc oxide.

**2. Oxidation of Copper:**



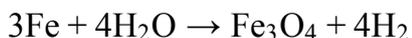
Copper reacts with oxygen to form copper(II) oxide.

**3. Reduction of Copper:**



Copper oxide is reduced to copper using hydrogen.

**4. Oxidation of Iron in Contact with Steam:**



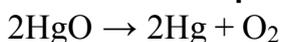
Iron reacts with steam to form magnetite and hydrogen.

**5. Disproportionation of Copper:**



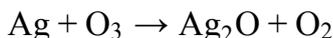
Copper(I) disproportionates to copper(II) and copper.

**6. Thermal Decomposition of Mercury(II) Oxide:**



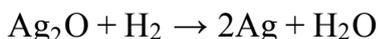
Mercury(II) oxide decomposes to mercury and oxygen on heating.

**7. Oxidation of Silver by Ozone:**



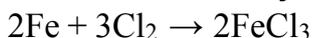
Silver is oxidized by ozone to form silver(I) oxide.

**8. Reduction of Silver(I) Oxide:**



Silver(I) oxide is reduced to silver using hydrogen.

**9. Oxidation of Iron by Chlorine:**



Iron reacts with chlorine to form iron(III) chloride.

**10. Reduction of Chromium(VI) to Chromium(III):**



Potassium dichromate is reduced to chromium(III) chloride in the presence of hydrochloric acid.

