

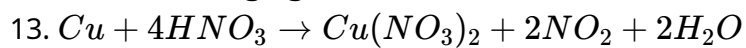
Saitechinfo NEET-JEE Academy

Here is the key to the questions for identifying the oxidizing and reducing agents:

- $Zn + Cu^{2+} \rightarrow Zn^{2+} + Cu$
 - **Oxidizing agent:** Cu^{2+}
 - **Reducing agent:** Zn
- $H_2 + F_2 \rightarrow 2HF$
 - **Oxidizing agent:** F_2
 - **Reducing agent:** H_2
- $2Fe^{3+} + 3Sn^{2+} \rightarrow 2Fe^{2+} + 3Sn^{4+}$
 - **Oxidizing agent:** Fe^{3+}
 - **Reducing agent:** Sn^{2+}
- $2H_2O_2 \rightarrow 2H_2O + O_2$
 - **Oxidizing agent:** H_2O_2
 - **Reducing agent:** H_2O_2 (acts as both)
- $Cl_2 + 2I^- \rightarrow 2Cl^- + I_2$
 - **Oxidizing agent:** Cl_2
 - **Reducing agent:** I^-
- $H_2SO_4 + 2HI \rightarrow I_2 + SO_2 + 2H_2O$
 - **Oxidizing agent:** H_2SO_4
 - **Reducing agent:** HI
- $MnO_4^- + 5Fe^{2+} + 8H^+ \rightarrow Mn^{2+} + 5Fe^{3+} + 4H_2O$
 - **Oxidizing agent:** MnO_4^-
 - **Reducing agent:** Fe^{2+}
- $HClO + H_2O_2 \rightarrow Cl^- + O_2 + H_2O$
 - **Oxidizing agent:** $HClO$
 - **Reducing agent:** H_2O_2
- $Cr_2O_7^{2-} + 6Cl^- + 14H^+ \rightarrow 2Cr^{3+} + 3Cl_2 + 7H_2O$
 - **Oxidizing agent:** $Cr_2O_7^{2-}$
 - **Reducing agent:** Cl^-
- $2NO + O_2 \rightarrow 2NO_2$
 - **Oxidizing agent:** O_2
 - **Reducing agent:** NO
- $2K + Br_2 \rightarrow 2KBr$
 - **Oxidizing agent:** Br_2
 - **Reducing agent:** K
- $H_2 + O_2 \rightarrow H_2O$

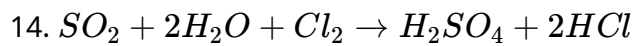
○ **Oxidizing agent:** O_2

○ **Reducing agent:** H_2



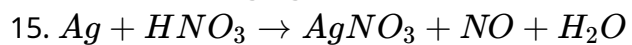
○ **Oxidizing agent:** HNO_3

○ **Reducing agent:** Cu



○ **Oxidizing agent:** Cl_2

○ **Reducing agent:** SO_2



○ **Oxidizing agent:** HNO_3

○ **Reducing agent:** Ag
