Hyphenated Term	Definition
State-function	A property whose value does not depend on the path taken to reach that specific value.
Heat-transfer	The movement of heat energy from one object to another.
Work-done	Energy transferred when a force moves an object.
First-law	A principle stating that energy cannot be created or destroyed, only transformed.
Second-law	A law stating that total entropy of an isolated system can only increase over time.
Open-system	A system that exchanges both energy and matter with its surroundings.
Closed-system	A system that exchanges energy but not matter with its surroundings.
Isolated-system	A system that does not exchange energy or matter with its surroundings.
Standard-enthalpy	The change in enthalpy when one mole of substance is formed under standard conditions.
Gibbs-free energy	A thermodynamic quantity that measures the useful work obtainable from a system.
Spontaneous-process	A process that occurs without needing to be driven by an external energy source.
Non-spontaneous	A process that requires external energy to proceed.
Equilibrium-state	A condition in which all competing influences are balanced in a system.
Reaction-enthalpy	The change in enthalpy when a reaction occurs at constant pressure.
System-boundary	The interface separating a system from its surroundings.
Path-function	A property that depends on the route taken between initial and final states.
Internal-energy	The total energy contained within a system.
Entropy-change	The change in entropy when a process occurs.
Reversible-process	A process that can be reversed without any net change in the system and surroundings.
Irreversible-process	A process that cannot return the system to its original state.
Heat-capacity	The amount of heat needed to increase the temperature of an object by a certain amount.
Molar-enthalpy	Enthalpy per mole of a substance.
Specific-heat	The heat required to raise the temperature of the unit mass of a substance by one degree.
Enthalpy-change	The total heat content change in a system at constant pressure.
Energy-conservation	The principle that the total energy of an isolated system remains constant.
Chemical-potential	The potential energy of molecules which drives chemical reactions.

Standard-state	A reference state for a substance used to calculate its properties under different conditions.
Partial-pressure	The pressure exerted by a single component in a mixture of gases.
Thermal-equilibrium	A condition where all parts of a system are at the same temperature.
Heat-of-reaction	The change in heat during a chemical reaction at constant pressure.
Equilibrium-constant	A number that expresses the relationship between products and reactants at equilibrium.