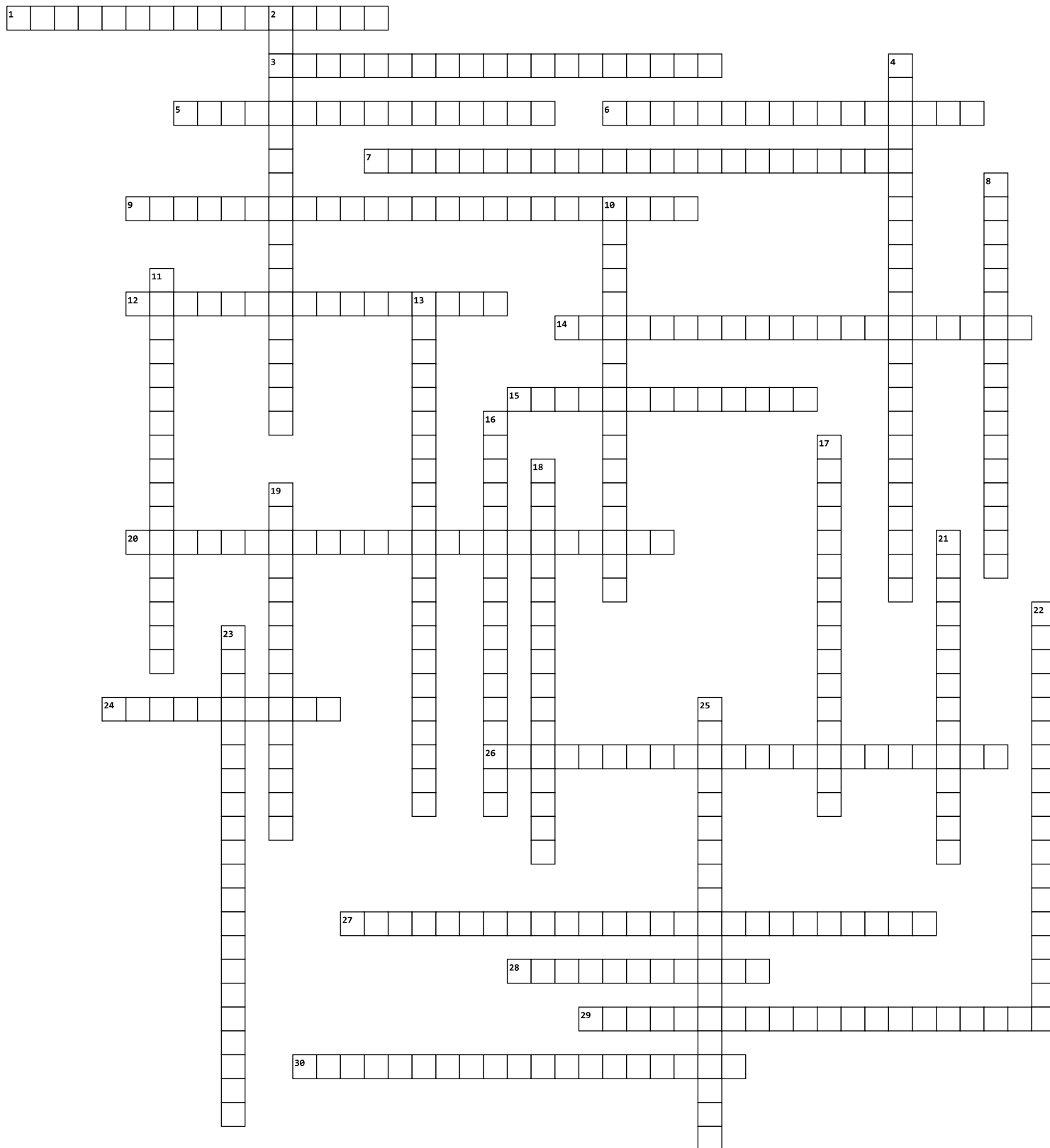


Chemical equilibrium



Across

1. The reaction proceeding in the direction of products formation.

Down

2. The potential energy of a species that influences its chemical reactions.

- 3.** The process of a system seeking the lowest energy configuration.
- 5.** The pressure exerted by a single gas in a mixture of gases.
- 6.** The influence of catalysts in altering equilibrium speed but not position.
- 7.** Equilibrium in reactions involving solid components.
- 9.** Equilibrium in reactions where all components are liquids.
- 12.** The reaction proceeding in the direction of reactants formation.
- 14.** A graph showing the progress of a reaction along its pathway.
- 15.** The ratio of moles of a component to the total moles in a mixture.
- 20.** An equilibrium involving reactants and products in the same phase.
- 24.** equation An equation relating equilibrium constant to temperature changes.
- 26.** Equilibrium involving ionization in aqueous solutions.
- 27.** An equilibrium involving reactants and products in different phases.
- 28.** law A law stating that reaction rate is proportional to active masses.
- 29.** A constant expressing the ratio of products to reactants at equilibrium.
- 30.** The composition of products and reactants at equilibrium.
- 4.** The stability of a system determined by energy levels at equilibrium.
- 8.** Equilibrium between different physical states of a substance.
- 10.** Equilibrium favoring reactant formation under given conditions.
- 11.** The ratio of product and reactant concentrations at any point in the reaction.
- 13.** The effect of temperature on the equilibrium constant.
- 16.** A change in equilibrium position due to external perturbations.
- 17.** Equilibrium favoring product formation under given conditions.
- 18.** The minimum energy required to initiate a chemical reaction.
- 19.** A situation where the reaction pathway determines product amounts.
- 21.** principle Principle stating how a system at equilibrium responds to changes in conditions.
- 22.** The equilibrium constant for sparingly soluble salts.
- 23.** Equilibrium in reactions where all components are gases.
- 25.** A state where the forward and reverse reactions occur at equal rates.